AQA

Please write clearly, in block capitals.

Centre number | | | | | | Candidate number | | | | |

Surname MODEL ANSWERS

Forename(s)

Candidate signature

A-level CHEMISTRY

Paper 3

Tuesday 27 June 2017 Morning Time allowed: 2 hours

Materials

For this paper you must have:
- the Periodic Table/Data Booklet, provided as an insert (enclosed)
- a ruler with millimetre measurements
- a calculator, which you are expected to use where appropriate.

Instructions
- Use black ink or black ball-point pen.
- Fill in the boxes at the top of the page.
- Answer **all** questions.
- You must answer the questions in the spaces provided.
  Do not write outside the box around each page or on blank pages.
- All working must be shown.
- Do all rough work in this book.
  Cross through any work you do not want to be marked.

Information
- The marks for questions are shown in brackets.
- The maximum mark for this paper is 90.

Advice
- You are advised to spend about 70 minutes on **Section A** and 50 minutes on **Section B**.

For Examiner's Use	
Question	Mark
1	
2	
3	
4	
Section B	
TOTAL	

Section A

Answer **all** questions in the spaces provided

0 1 Anhydrous magnesium chloride, $MgCl_2$, can absorb water to form the hydrated salt $MgCl_2.4H_2O$

$$MgCl_2(s) + 4H_2O(l) \rightarrow MgCl_2.4H_2O(s)$$

0 1 . 1 Suggest **one** reason why the enthalpy change for this reaction cannot be determined directly by calorimetry.

[1 mark]

Its not possible to prevent some dissolving

0 1 . 2 Some enthalpies of solution are shown in **Table 1**.

Table 1

Salt	Enthalpy of solution / kJ mol^{-1}
$MgCl_2(s)$	−155
$MgCl_2.4H_2O(s)$	−39

Calculate the enthalpy change for the absorption of water by $MgCl_2(s)$ to form $MgCl_2.4H_2O(s)$.

[2 marks]

anhydrous $MgCl_2(s) + aq \xrightarrow{\Delta H_r^\ominus} MgCl_2.4H_2O + aq$ hydrated

$\Delta H_{sol}^\ominus = -155$ ↘ ↙ $\Delta H_{sol}^\ominus = -39$

$MgCl_2(aq)$ M1

$\Delta H_r^\ominus = -155 - (-39)$

Enthalpy change -116 kJ mol^{-1} M2

0 1 . 3 Describe how you would carry out an experiment to determine the enthalpy of solution of anhydrous magnesium chloride.

LEVELLED You should use about 0.8 g of anhydrous magnesium chloride.

Explain how your results could be used to calculate the enthalpy of solution.

[6 marks]

Method:

- Measure 25 cm^3 deionised water using a measuring cylinder.
- Place into a polystyrene cup with a lid
- Record the initial temperature at one minute intervals for 3 minutes.
- Weigh 0.8g anhydrous magnesium chloride into a stoppered weighing bottle.
- At the 4th minute add the $MgCl_2$ to the water and stir.
- Re-weigh the weighing bottle and record mass.
- Continue to stir the solution.
- At the 5th minute and every minute up to 15 minutes record the temperature.

Results:

- Plot temperature v. time
- Extrapolate the lines before and after addition of $MgCl_2$ to find the initial and final temperature.
- Calculate ΔT from the graph readings
- Calculate enthalpy $q = mc\Delta T$
- Calculate moles $MgCl_2$ used $n = m/M_r$
- Calculate enthalpy change $\Delta H_{sol} = q/n$

0 1 . 4 Anhydrous magnesium chloride can be formed by direct reaction between its elements.

$$Mg(s) + Cl_2(g) \rightarrow MgCl_2(s)$$

The free-energy change, ΔG, for this reaction varies with temperature as shown in **Table 2**.

Table 2

T / K	ΔG / kJ mol^{-1}
298	−592.5
288	−594.2
273	−596.7
260	−598.8
240	−602.2

Use these data to plot a graph of free-energy change against temperature on the grid opposite.

Calculate the gradient of the line on your graph and hence calculate the entropy change, ΔS, in J K^{-1} mol^{-1}, for the formation of anhydrous magnesium chloride from its elements.

Show your working.

[5 marks]

Gradient $= \frac{\Delta(\Delta G)}{\Delta T} = \frac{-592.5 - (-602.2)}{298 - 240} = \frac{9.7}{58}$ M3

$= 0.167$ kJ mol^{-1}

$\Delta G = \Delta H - T\Delta S$ Re-arrange to give format

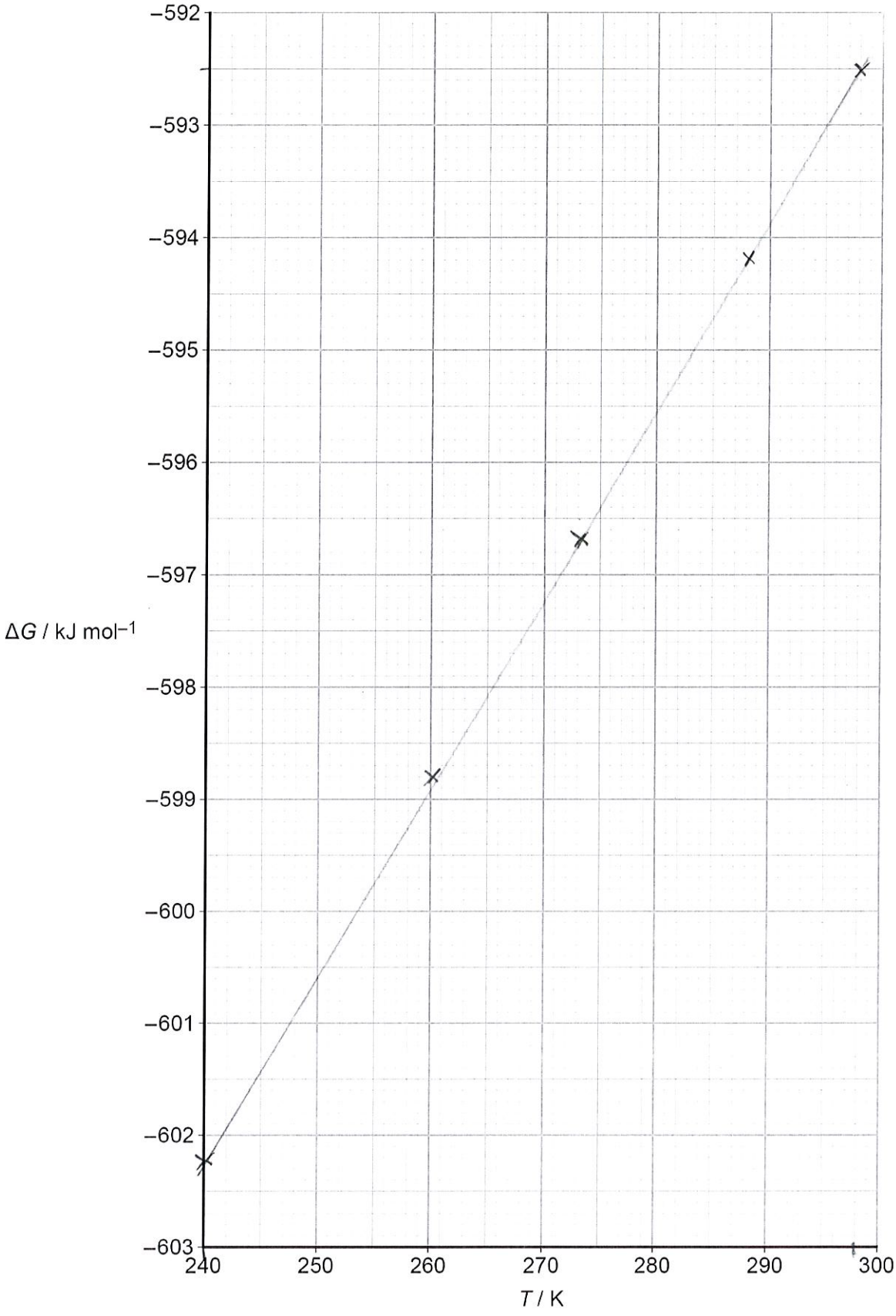
$\Delta G = -\Delta S T + \Delta H$ $y = mx + c$

So gradient $= -\Delta S$ J K^{-1} mol^{-1}

$\therefore \Delta S = -0.167 \times 10^3 = -167$ J K^{-1} mol^{-1} M5

M4 (unit conversion)

(Allow −163 to −171)



M1
5 correct points

M2
line (not curved doubled or linked

ΔS _____ $JK^{-1}mol^{-1}$

14

0 2 Concentrated sulfuric acid reacts with alkenes, alcohols and sodium halides.

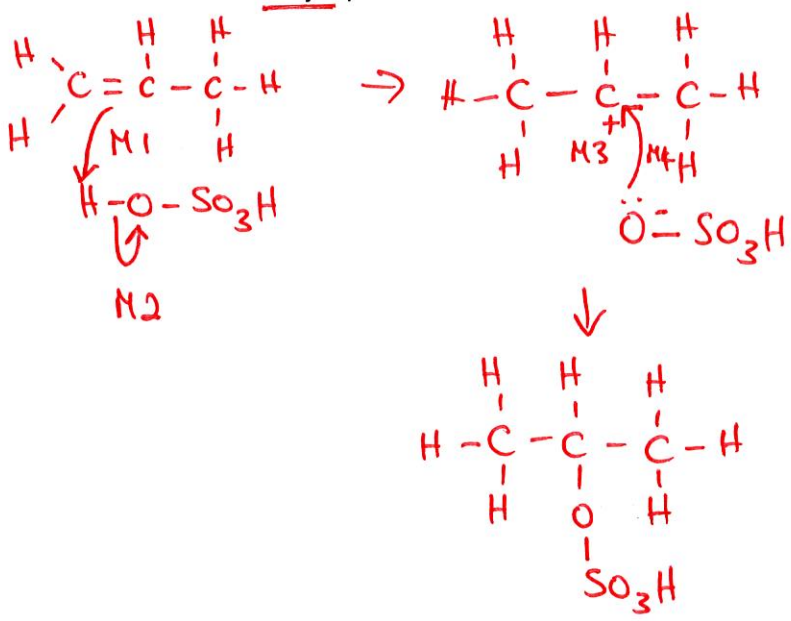
0 2 . 1 Name the mechanism for the reaction of concentrated sulfuric acid with an alkene.

[1 mark]

Electrophilic addition

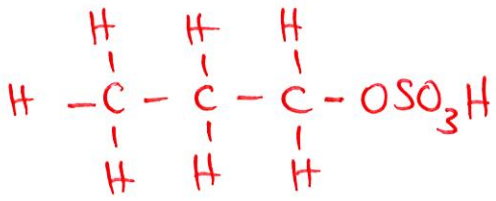
0 2 . 2 Outline the mechanism for the reaction of concentrated sulfuric acid with propene to show the formation of the major product.

[4 marks]



0 2 . 3 Draw the structure of the minor product of the reaction between concentrated sulfuric acid and propene.

[1 mark]



0 2 . 4 Explain why the product shown in your answer to Question **2.2** is the major product.

[2 marks]

- Major product is formed via the more stable secondary carbocation M1
- Due to the positive inductive effect of 2 alkyl groups M2

0 2 . 5 Butan-2-ol reacts with concentrated sulfuric acid to form a mixture of three isomeric alkenes. Two of the alkenes are stereoisomers.

Draw the skeletal formula of each of the three isomeric alkenes formed by the reaction of butan-2-ol with concentrated sulfuric acid.

Give the full IUPAC name of each isomer.

[3 marks]

Skeletal formula	Name
	but-1-ene
	E-but-2-ene
	Z-but-2-ene

0 2 . 6 A by-product of the reaction of butan-2-ol with concentrated sulfuric acid has the molecular formula C_4H_8O

Name this by-product, identify the role of the sulfuric acid in its formation and suggest the name of a method that could be used to separate the products of this reaction.

[3 marks]

By-product butanone

Role of sulfuric acid oxidising agent

Name of separation method

distillation

0 2 . 7 Concentrated sulfuric acid reacts with solid sodium chloride.

Give the observation you would make in this reaction.
State the role of the sulfuric acid.

[2 marks]

Observation with sodium chloride misty / steamy fumes

Role of sulfuric acid acid / proton donor

0 2 . 8 Concentrated sulfuric acid reacts with solid sodium iodide, to produce several products.

Observations made during this reaction include the formation of a black solid, a yellow solid and a gas with the smell of bad eggs.

Identify the product responsible for each observation.

[3 marks]

Black solid iodine / I_2

Yellow solid sulphur / S / S_8

Gas hydrogen sulphide / H_2S

19

0 3 . 3 Suggest why an excess of sodium hydroxide is used.

[1 mark]

To ensure all the ethyl benzoate is completely reacted

0 3 . 4 Suggest why an electric heater is used rather than a Bunsen burner in this hydrolysis.

[1 mark]

To prevent ignition of any flammable vapours

0 3 . 5 State why reflux is used in this hydrolysis.

[1 mark]

It allows reactant vapours to be returned to the reaction mixture

or Doesn't allow reactant vapours to escape

0 3 . 6 Write an equation for the reaction between sodium benzoate and hydrochloric acid.

[1 mark]

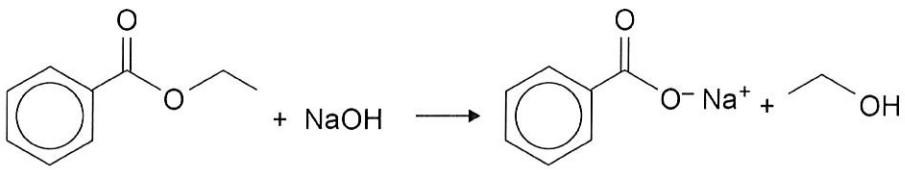
$C_6H_5COONa + HCl \rightarrow C_6H_5COOH + NaCl$

0 3 . 7 Suggest why sodium benzoate is soluble in cold water but benzoic acid is insoluble in cold water.

[2 marks]

- Sodium benzoate is ionic M1
- Benzoic acid contains a non-polar benzene ring which makes it insoluble even though the COOH is able to form hydrogen bonds

0 3 Benzoic acid can be prepared from ethyl benzoate.
Ethyl benzoate is first hydrolysed in alkaline conditions as shown:



A student used the following method.

Add 5.0 cm^3 of ethyl benzoate (density = 1.05 g cm^{-3}, M_r = 150) to 30.0 cm^3 of aqueous 2 mol dm^{-3} sodium hydroxide in a round-bottomed flask.

Add a few anti-bumping granules and attach a condenser to the flask. Heat the mixture under reflux for half an hour. Allow the mixture to cool to room temperature.

Pour 50.0 cm^3 of 2 mol dm^{-3} hydrochloric acid into the cooled mixture.

Filter off the precipitate of benzoic acid under reduced pressure.

0 3 . 1 Suggest how the anti-bumping granules prevent bumping during reflux.

[1 mark]

Allow small bubbles to form
or Prevent large bubbles forming

0 3 . 2 Show, by calculation, that an excess of sodium hydroxide is used in this reaction.

[2 marks]

$n\,NaOH = 2 \times 30 \times 10^{-3} = 0.06$ mol M1

$n = \frac{c}{v}$

n ethyl benzoate $= \frac{m}{Mr}$

m ethyl benzoate $= 1.05 \times 5 = 5.25$ g

$m = dv$

$n = \frac{5.25}{150} = 0.035$ mol M2

Question 3 continues on the next page

0 3 . 8 After the solid benzoic acid has been filtered off, it can be purified.

Describe the method that the student should use to purify the benzoic acid.

[6 marks]

- Dissolve crude benzoic acid in hot solvent M1
- Use minimum volume M2
- filter whilst hot to remove insoluble impurities M3
- Cool to recrystalise M4
- Filter under reduced pressure M5
- Wash with cold solvent and dry M6

Question 3 continues on the next page

0 3 . 9 In a similar experiment, another student used 0.040 mol of ethyl benzoate and obtained 5.12 g of benzoic acid.

1:1 mole ratio from equation

Calculate the percentage yield of benzoic acid.

Suggest why the yield is not 100%.

Mr benzoic acid = 122

[3 marks]

n benzoic acid = $\frac{5.12}{122}$ = 0.042 (produced more) M1

$\frac{0.042}{0.040} \times 100 = 104.9\%$

Percentage yield 105 % M2

Suggestion Impurities present in benzoic acid (water present) M3

18

0 4 . 3 Calculate the concentration of HX in the original solution.

[2 marks]

Neutralisation = 24.0 cm^3 NaOH (from graph)

$n\ NaOH = 0.1 \times 24 \times 10^{-3} = 2.4 \times 10^{-3}$ M1

$n\ HX = 2.4 \times 10^{-3}$

$c\ HX = \frac{2.4 \times 10^{-3}}{25 \times 10^{-3}}$

Concentration 0.0960 mol dm^{-3} M2

0 4 . 4 Calculate the pH of the solution of HX before the addition of any sodium hydroxide.

(If you were unable to calculate a value for the concentration of HX in Question **4.3** you should use a value of 0.600 mol dm^{-3} in this calculation. This is **not** the correct value.)

[2 marks]

$ka = \frac{[H^+]^2}{[HX]}$

$[H^+] = \sqrt{ka\,[HX]}$ M1

$= \sqrt{2.62 \times 10^{-5} \times 0.096}$

$= 1.586 \times 10^{-3}$

$pH = -\log [H^+]$

pH of HX 2.80 (2d.p.) M2

0 4 . 5 Calculate the pH of the solution when half of the acid has reacted.

[1 mark]

At half-neutralisation pka = pH

$pka = -\log 2.62 \times 10^{-5}$

pH of solution 4.58

0 4 . 6 Plot your answers to Questions **4.4** and **4.5** on the grid in **Figure 1**.
Use these points to sketch the missing part of the curve between 0 and 20 cm^3 of NaOH solution added.

[2 marks]

9

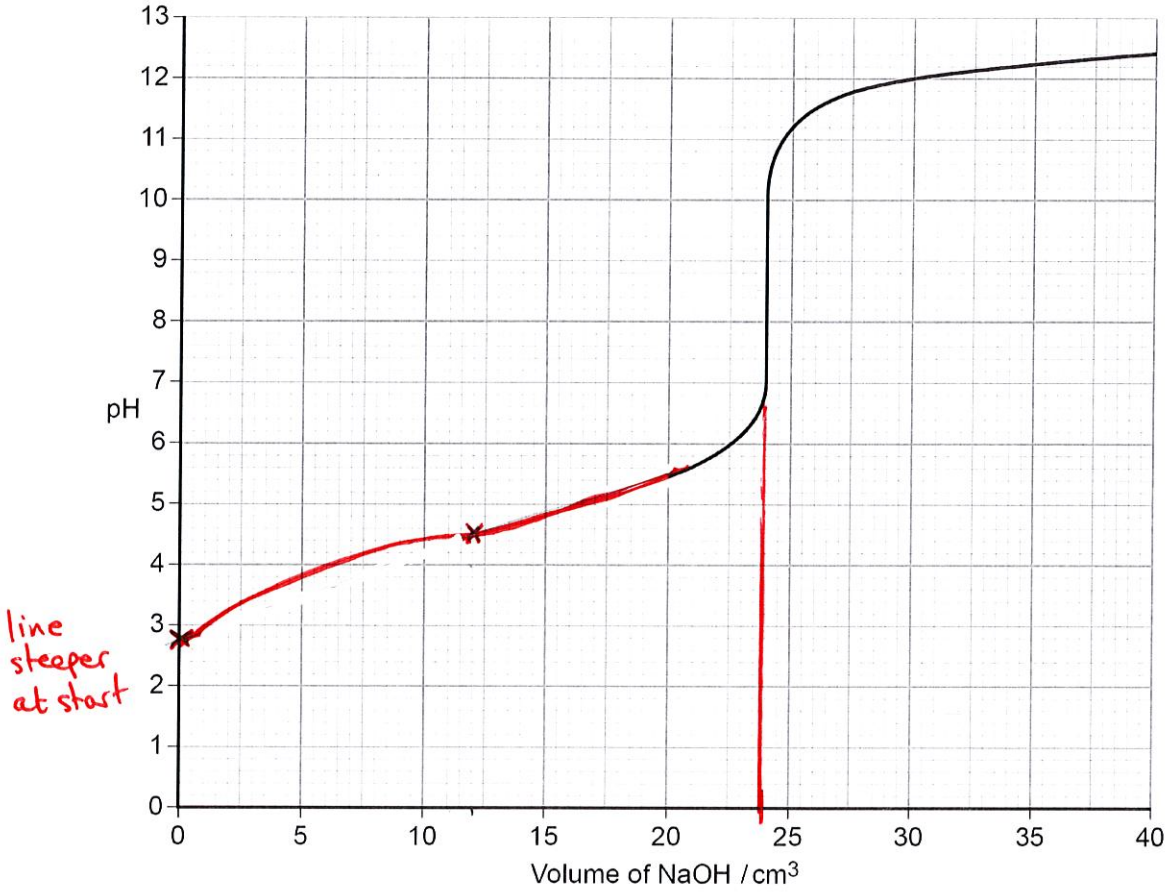
0 4 A 0.100 mol dm^{-3} solution of sodium hydroxide was gradually added to 25.0 cm^3 of a solution of a weak acid, HX, in the presence of a suitable indicator.

A graph was plotted of pH against the volume of sodium hydroxide solution, as shown in **Figure 1**.

The first pH reading was taken after 20.0 cm^3 of sodium hydroxide solution had been added.

The acid dissociation constant of HX, K_a, = 2.62×10^{-5} mol dm^{-3}

Figure 1



0 4 . 1 The pH range of an indicator is the range over which it changes colour.

Suggest the pH range of a suitable indicator for this titration.

[1 mark]

7 - 10.2 (any 2 values in this range)

0 4 . 2 Give the expression for the acid dissociation constant of HX.

$HX \rightarrow H^+ + X^-$

[1 mark]

$K_a = \dfrac{[H^+][X^-]}{[HX]}$

Section B

Answer **all** questions in the spaces provided

Only **one** answer per question is allowed.
For each answer completely fill in the circle alongside the appropriate answer.

CORRECT METHOD ● WRONG METHODS ⊗ ⊙ ≈ √

If you want to change your answer you must cross out your original answer as shown. ⊗

If you wish to return to an answer previously crossed out, ring the answer you now wish to select as shown. ⊗

You may do your working in the blank space around each question but this will not be marked.
Do **not** use additional sheets for this working.

0 5 Which compound has the highest boiling point?

[1 mark]

A $CH_3CH_2CH_2OH$ hydrogen bonding ●

B CH_3CH_2CHO ○

C CH_3COCH_3 ○

D CH_3COOCH_3 ○

0 6 Which is the correct order of melting points of these Period 3 elements?

[1 mark]

A phosphorus > sulfur > chlorine > argon ○

B argon > chlorine > phosphorus > sulfur ○

C sulfur > phosphorus > chlorine > argon ● S_8 P_4 Cl_2 Ar

D chlorine > phosphorus > sulfur > argon ○

Turn over for the next question

0 7 Which is **not** a correct statement?

[1 mark]

A Transition metals form coloured ions and complexes ○

B Transition metals display variable oxidation states ○

C A ligand accepts a pair of electrons from a transition metal ●

donates to

D A complex is a central metal atom or ion surrounded by ligands ○

0 8 The table shows possible conditions and products for the cracking of alkanes.

Which row is correct?

[1 mark]

	Type of cracking	Conditions	Products	
A	Thermal	High pressure High temperature	Mainly alkanes	○
B	Thermal	Slight pressure High temperature	Mainly alkenes	○
C	Catalytic	Slight pressure High temperature	Mainly branched alkanes and aromatics	●
D	Catalytic	High pressure High temperature	Mainly branched alkanes and aromatics	○

0 9 2,4,6-Trichlorophenol is a weak monoprotic acid, with $K_a = 2.51 \times 10^{-8}$ mol dm^{-3} at 298 K.

What is the concentration, in mol dm^{-3}, of hydrogen ions in a 2.00×10^{-3} mol dm^{-3} solution of 2,4,6-trichlorophenol at 298 K?

[1 mark]

A 5.02×10^{-11} ○

B 7.09×10^{-6} ●

C 1.26×10^{-5} ○

D 3.54×10^{-3} ○

$[H^+] = \sqrt{K_a [HA]}$

$= \sqrt{2.51 \times 10^{-8} \times 2 \times 10^{-3}}$

$= 7.09 \times 10^{-6}$

1 0 What is the pH of a 0.46 mol dm^{-3} solution of potassium hydroxide at 298 K?

(K_w = 1.0 × 10^{-14} mol^2 dm^{-6} at 298 K)

[1 mark]

A 0.34 $kw = [H^+][OH^-]$

B 13.66 $[H^+] = kw / [OH^-]$

C 13.96 $= 1 \times 10^{-14} / 0.46 = 2.17 \times 10^{-14}$

D 14.34 $pH = -\log [H^+]$

1 1 What is the mass, in mg, of carbon formed when 3.0 × 10^{-3} mol of propene undergoes incomplete combustion?

$$2C_3H_6 + 3O_2 \rightarrow 6C + 6H_2O$$

[1 mark]

A 9.0 × 10^{-3} $nC = 3.0 \times 10^{-3} \times 3$

$= 9.0 \times 10^{-3}$

B 3.6 × 10^{-2} $mC = 9.0 \times 10^{-3} \times 12$

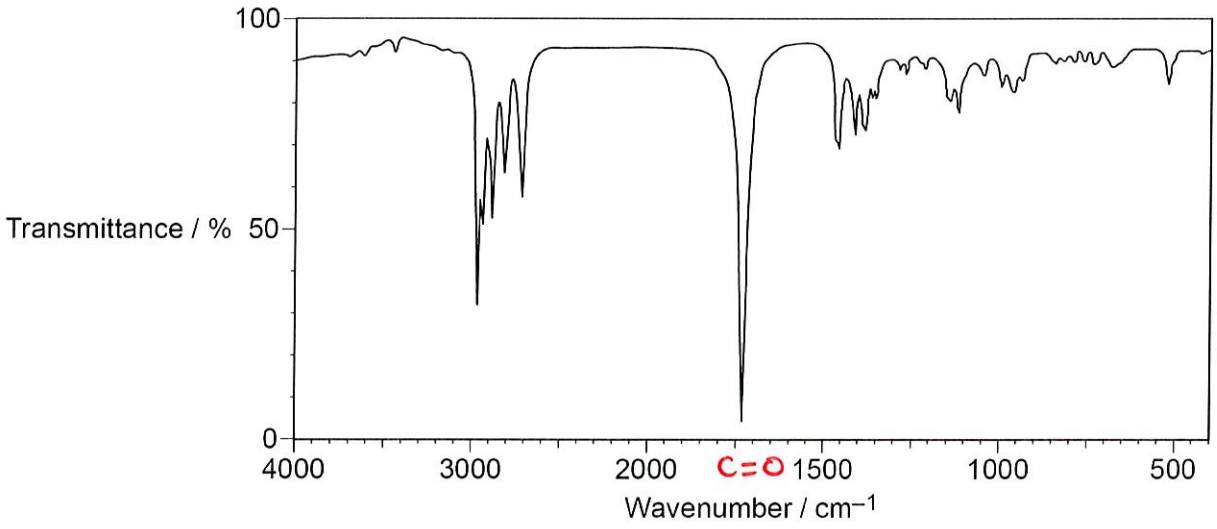
C 1.08 × 10^2 $= 0.108g \times 10^3$

D 2.16 × 10^2 $= 108\ mg$

Turn over for the next question

1 2 Which compound gives this infrared spectrum?

[1 mark]



A 1-bromobutane

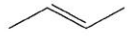
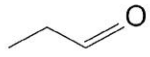
B butan-1-ol

C butanal

D butanoic acid

1 3 Which pair of compounds does **not** form a racemic mixture when the compounds react?

[1 mark]

A		+ HCl
B		+ HCN
C		+ HCl
D		+ HCN

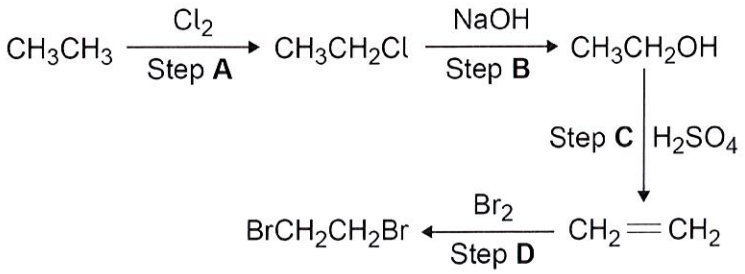
A ⬭

B ⬭

C ⬭

D ⬬

1 4 The reaction sequence shows how CH_3CH_3 can be converted into $BrCH_2CH_2Br$



Which step occurs by nucleophilic substitution?

[1 mark]

A Step **A** ⬭

B Step **B** ⬬

C Step **C** ⬭

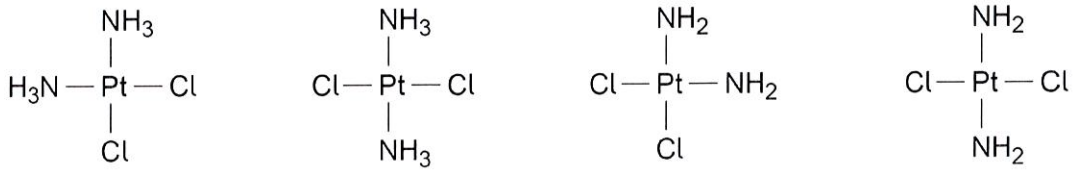
D Step **D** ⬭

Turn over for the next question

1 5 Cisplatin is an anti-cancer drug.

Which structure represents a stereoisomer of cisplatin?

[1 mark]



A **B** **C** **D**

transplatin

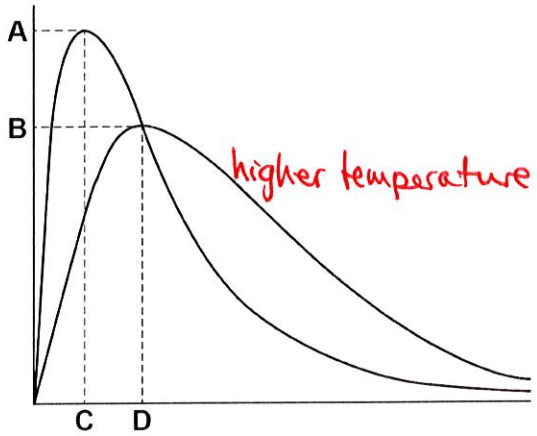
A ○

B ●

C ○

D ○

1 6 The diagram shows the Maxwell–Boltzmann distribution of molecular energies in a gas at two different temperatures.



Which letter represents the most probable energy of the molecules at the higher temperature?

[1 mark]

A ○

B ○

C ○

D ●

1 7 V_2O_5 can be used as a catalyst in the Contact Process.

oxygen added

Which is a step in the Contact Process in which the vanadium is oxidised?

loses e^-
ON increases

[1 mark]

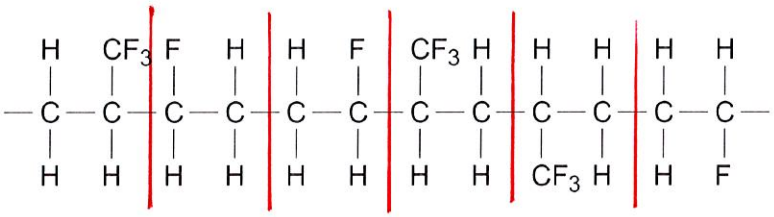
A $SO_2 + V_2O_5 \rightarrow SO_3 + 2VO_2$ ○

B $SO_3 + 2VO_2 \rightarrow SO_2 + V_2O_5$ ○

C $2VO_2 + \frac{1}{2}O_2 \rightarrow V_2O_5$ ●

D $V_2O_5 \rightarrow 2VO_2 + \frac{1}{2}O_2$ ○

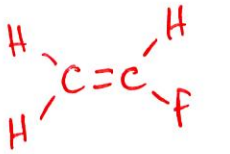
1 8 This structure shows a section of a polymer chain formed from the random polymerisation of two different monomers.



Which pair of monomers could produce this polymer?

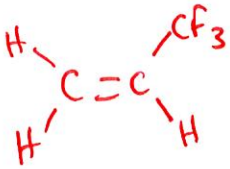
[1 mark]

A $CH_2=CHF$ and $CH_2=CHCF_3$ ●



B $CH_2=CH_2$ and $CHF=CHCF_3$ ○

C $CH_2=CH_2$ and $CH_2=CHCF_3$ ○



D $CH_2=CHF$ and $CHCF_3=CHF$ ○

Turn over for the next question

19 The equation for the reaction between zinc and hydrochloric acid is

$$Zn + 2HCl \rightarrow ZnCl_2 + H_2$$

What is the minimum mass, in mg, of zinc (A_r = 65.4) needed to react with 50.0 cm^3 of 1.68 mol dm^{-3} hydrochloric acid?

[1 mark]

A 2.75 $n_{HCl} = 1.68 \times 50 \times 10^{-3} = 0.084$

B 5.49 $n_{Zn} = 0.084/2 = 0.042$

C 2.75×10^3 $m_{Zn} = 0.042 \times 65.4$

D 5.49×10^3 $= 2.75\,g$

20 An equilibrium mixture is prepared in a container of fixed volume.

$$CO(g) + Cl_2(g) \rightleftharpoons COCl_2(g) \qquad \Delta H = -108\ kJ\ mol^{-1}$$

exothermic

The temperature of this mixture is decreased and the mixture is allowed to reach a new equilibrium.

Which is greater for the new equilibrium than for the original equilibrium?

[1 mark]

A The mole fraction of carbon monoxide smaller

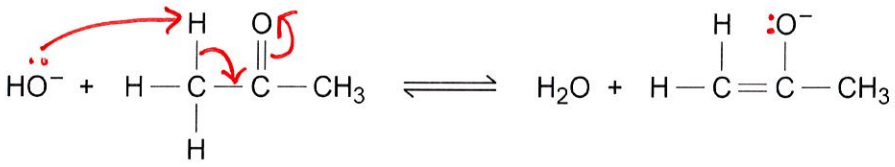
B The partial pressure of chlorine smaller

C The total pressure of the mixture same

D The value of the equilibrium constant, K_p greater

eqm more RHS

2 1 In concentrated alkali, propanone reacts with hydroxide ions to form an equilibrium mixture as shown.

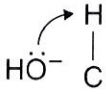


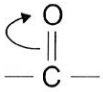
Which curly arrow does **not** appear in the mechanism of this reaction?

[1 mark]

A 

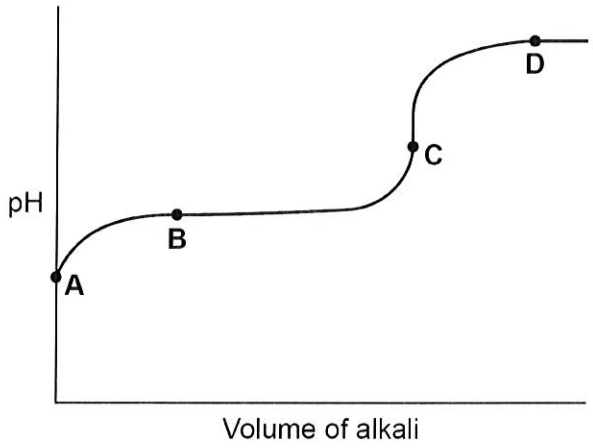
B 

C 

D 

Turn over for the next question

2 2 The diagram shows a pH curve produced by adding a strong alkali to a weak acid.



Which point on the curve represents a solution that can act as a buffer?

[1 mark]

A ○

B half-neutralisation ●

C ○

D ○

2 3 Which alcohol could **not** be produced by the reduction of an aldehyde or a ketone?

[1 mark]

A 2,2-dimethylpropan-1-ol 1° ○

B 2-methylbutan-2-ol 3° ●

C 3-methylbutan-2-ol 2° ○

D pentan-3-ol 1° ○

2 4 Which compound does **not** show stereoisomerism?

[1 mark]

A 1,2-dichloropropene E-Z ○

B 1,2-dichloropropane chiral carbon ○

C 1,3-dichloropropene E-Z ○

D 1,3-dichloropropane ●

2 5 Which compound can form a polymer without needing another reagent?

[1 mark]

A $HOCH_2CH_2OH$ ○

B $HOOCCH_2CH_2COOH$ ○

C $HOCH_2CH_2COCl$ 2 functional groups plus highly reactive acyl chloride ●

D $ClCH_2CH_2COOH$ needs catalyst ○

2 6 A solution of lead(II) chloride (M_r = 278.2) contains 1.08 g of $PbCl_2$ in 100 cm^3 of solution. In this solution, the lead(II) chloride is fully dissociated into ions.

What is the concentration of chloride ions in this solution?

[1 mark]

$PbCl_2 \rightarrow Pb^{2+} + 2Cl^-$

A 3.88×10^{-3} mol dm^{-3} $n PbCl_2 = \frac{1.08}{278.2} = 3.88 \times 10^{-2}$ ○

B 7.76×10^{-3} mol dm^{-3} $nCl^- = 3.88 \times 10^{-2} \times 2$ ○

C 3.88×10^{-2} mol dm^{-3} $= 7.76 \times 10^{-3}$ ○

D 7.76×10^{-2} mol dm^{-3} $cCl^- = \frac{7.76 \times 10^{-3}}{100 \times 10^{-3}}$ ●

$= 7.76 \times 10^{-2}$

Turn over for the next question

2 7 The rate equation for the acid-catalysed reaction between iodine and propanone is:

$$\text{rate} = k\,[H^+]\,[C_3H_6O]$$

The rate of reaction was measured for a mixture of iodine, propanone and sulfuric acid at pH = 0.70

In a second mixture the concentration of the sulfuric acid was different but the concentrations of iodine and propanone were unchanged. The new rate of reaction was a quarter of the original rate.

What was the pH of the second mixture?

[1 mark]

A 1.00 ○

B 1.30 ●

C 1.40 ○

D 2.80 ○

2 8 A 385 cm^3 sample of carbon dioxide at 100 kPa and 25 °C was mixed with 2.89×10^{-2} mol of argon. The gas constant, $R = 8.31\ J\ K^{-1}\ mol^{-1}$

What is the mole fraction of carbon dioxide in the mixture?

[1 mark]

$pV = nRT \quad n = \frac{pV}{RT}$

$\frac{100 \times 10^3 \times 385 \times 10^{-6}}{8.31 \times 298} = 1.55 \times 10^{-2}$

mole fraction

$\frac{1.55 \times 10^{-2}}{(1.55 \times 10^{-2} + 2.89 \times 10^{-2})}$

A 0.35 ●

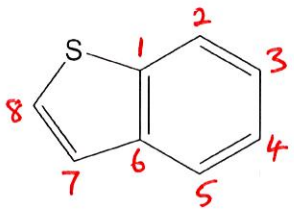
B 0.46 ○

C 0.54 ○

D 0.65 ○

2 9 How many peaks does this compound have in its ^{13}C spectrum?

[1 mark]



A 5 ○

B 6 ○

C 7 ○

D 8 ●

3 0 A student is provided with 5.00 cm^3 of 1.00 mol dm^{-3} ammonia solution. The student was asked to prepare an ammonia solution with a concentration of 0.050 mol dm^{-3}

What volume of water should the student add?

moles NH_3 same in each solution so $n_1 = n_2$

[1 mark]

A 45.0 cm^3 $C_1V_1 = C_2V_2$

B 95.0 cm^3 $1.00 \times 5.00 \times 10^{-3} = 0.050\ V_2$

C 100 cm^3 $V_2 = \frac{5 \times 10^{-3}}{0.05} = 0.1\ dm^3$ ●

D 995 cm^3 ○

3 1 A solution absorbs light with wavelengths corresponding to red, yellow and green light.

Which ion is most likely to be in the solution? not blue

[1 mark]

A $Cr_2O_7^{2-}(aq)$ ○

B $Fe^{2+}(aq)$ ○

C $Fe^{3+}(aq)$ ○

D $Cu^{2+}(aq)$ look blue ●

Turn over for the next question

3 2 A reaction is exothermic and has a negative entropy change.

Which statement is correct?

feasible if $\Delta G \leq 0$

[1 mark]

A The reaction is always feasible ○

B The reaction is feasible above a certain temperature ○

C The reaction is feasible below a certain temperature ●

D The reaction is never feasible ○

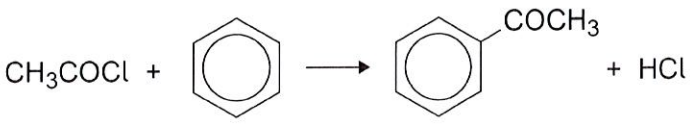
$\Delta G = \Delta H - T\Delta S$

$(-ve) - T(-ve)$

for ΔG to be ≤ 0

T must be small

3 3 Phenylethanone can be prepared by the reaction between ethanoyl chloride and benzene.



In a preparation, with an excess of benzene, the mass of ethanoyl chloride (M_r = 78.5) used was 5.7×10^{-2} kg.

The percentage yield of phenylethanone was 62%.

What mass, in grams, of phenylethanone was produced?

[1 mark]

A 35 g

B 54 g

C 87 g

D 102 g

$n\ CH_3COCl = \frac{5.7 \times 10^{-2} \times 10^3}{78.5} = 0.726$

$n\ C_6H_5COCH_3 = 0.726$

$Mr\ C_6H_5COCH_3 = 120$

$m\ C_6H_5COCH_3 = 0.726 \times 120$

$= 87.12\ g$

$87.12 \times 62\% = 54\ g$

3 4 130 cm^3 of oxygen and 40 cm^3 of nitrogen, each at 298 K and 100 kPa, were placed into an evacuated flask of volume 0.50 dm^3.

What is the pressure of the gas mixture in the flask at 298 K?

[1 mark]

130 + 40 / 500 = 0.34

100 × 0.34

A 294 kPa ○

B 68.0 kPa ○

C 34.0 kPa ●

D 13.7 kPa ○

30

END OF QUESTIONS