Gateways School

**Alkenes & Alcohols**

**Revision PPQ**

40 marks

**Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_**

**Q1.**

Propene reacts with bromine by a mechanism known as electrophilic addition.

(a)     Explain what is meant by the term *electrophile* and by the term *addition*.

*Electrophile* \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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*Addition* \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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**(2)**

(b)     Explain why bromine, a non-polar molecule, is able to react with propene.

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**(2)**

(c)     Outline the mechanism for the electrophilic addition of bromine to propene. Give the name of the product formed.

*Mechanism*

*Name of product* \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**(5)**

(d)     The polymerisation of propene to form poly(propene) is an important industrial process.

Name the type of polymerisation involved.

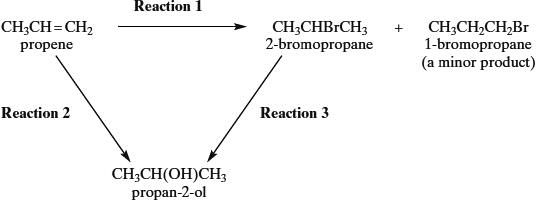
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**(1)**

**(Total 10 marks)**

**Q2.**

Consider the following reaction scheme.



(a)     (i)      Name the mechanism for **Reaction 1**.

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(ii)     Explain why 1-bromopropane is only a minor product in **Reaction 1**.

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**(3)**

(b)     Give a suitable reagent and state the essential conditions required for **Reaction 3**.

*Reagent* \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

*Conditions* \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**(2)**

(c)     The reagent used for **Reaction 3** can also be used to convert 2-bromopropane into propene. State the different conditions needed for this reaction.

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**(1)**

(d)     **Reaction 2** proceeds in two stages.

*Stage 1*      CH3CH=CH2 + H2SO4  →  CH3CH(OSO2OH)CH3

*Stage 2*      CH3CH(OSO2OH)CH3 + H2O  →  CH3CH(OH)CH3 + H2SO4

(i)      Name the class of alcohols to which propan-2-ol belongs.

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(ii)     Outline a mechanism for Stage 1 of **Reaction 2**, using concentrated sulphuric acid.

(iii)     State the overall role of the sulphuric acid in **Reaction 2**.

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**(6)**

**(Total 12 marks)**

**Q3.**

The following instructions are from an experimental procedure for the preparation of cyclohexene from cyclohexanol and concentrated phosphoric acid.   
Read these instructions and answer the questions that follow.

1        Place 25 cm3 of cyclohexanol into a round-bottomed flask with some porous pot to act as anti-bumping granules. Add 10 cm3 of concentrated phosphoric acid carefully while shaking the flask. Cool the flask under the tap if it gets too hot.   
Make sure the reagents are thoroughly mixed.

2        Set up an apparatus for simple distillation using this flask.

3        Warm the flask, gently at first, for about 15 minutes. Then increase the heating so that cyclohexene begins to distil over. Collect the fraction that distils below 95 °C.

(a)     State the purpose of the anti-bumping granules.

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**(1)**

(b)     Name the part of the distillation apparatus where cyclohexene vapour is changed back into a liquid.   
Draw a simple diagram of this part of the apparatus.

Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Diagram

**(2)**

**(Total 3 marks)**

**Q4.**

Some alcohols can be oxidised by an acidified solution of potassium dichromate(Vl).   
Aldehydes can be oxidised by Tollens’ reagent or by Fehling’s solution.

An unknown pure liquid **A** contains only a single alcohol.   
Outline a simple procedure to allow you to determine whether **A** is a primary, a secondary or a tertiary alcohol.

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**(Total 3 marks)**

**Q5.**

Glucose can be used as a source of ethanol. Ethanol can be burned as a fuel or can be converted into ethene.

C6H12O6   →   CH3CH2OH   →   H2C=CH2

glucose            ethanol            ethene

(a)     Name the types of reaction illustrated by the two reactions above.

*Glucose to ethanol* \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

*Ethanol to ethene* \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**(2)**

(b)     (i)      State what must be added to an aqueous solution of glucose so that ethanol is formed.

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(ii)     Identify a suitable catalyst for the conversion of ethanol into ethene.

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**(2)**

(c)     (i)      State the class of alcohols to which ethanol belongs.

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(ii)     Give **one** advantage of using ethanol as a fuel compared with using a petroleum fraction.

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**(2)**

(d)     Most of the ethene used by industry is produced when ethane is heated to 900°C in the absence of air. Write an equation for this reaction.

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**(1)**

(e)     Name the type of polymerisation which occurs when ethene is converted into poly(ethene).

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**(1)**

**(Total 8 marks)**

**Q6.**

The correct name for the alkene monomer which forms the polymer shown below is



**A**       2-methyl-3-ethylpropene

**B**       2-methylpent-2-ene

**C**       2-methylpent-3-ene

**D**       4-methylpent-2-ene

**(Total 1 mark)**

**Q7.**

The structure of the molecule of methyl 2-methylpropenoate is shown below.



Which one of the following statements concerning this compound is **not** true?

**A**       It displays geometrical isomerism.

**B**       It forms an addition polymer.

**C**       It undergoes reduction.

**D**       It decolourises bromine.

**(Total 1 mark)**

**Q8.**

Which one of the following **cannot** be produced by oxidation of propan-l-ol?

**A**       carbon dioxide

**B**       propanone

**C**       propanal

**D**       propanoic acid

**(Total 1 mark)**

**Q9.**

Which one of the following isomers is not oxidised under mild reaction conditions?

**A**       (CH3)2CHCH(OH)COCH3

**B**       (CH3)2C(OH)CH2COCH3

**C**       (CH3)2CHCH(OH)CH2CHO

**D**       (CH3)2C(OH)CH2CH2CHO

**(Total 1 mark)**