Gateways School

**Period 3 & Reactions of aqueous ions**

**Revision PPQ**

39 marks

**Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_**

**Q1.**

(a)     **P** and **Q** are oxides of Period 3 elements.

Oxide **P** is a solid with a high melting point. It does not conduct electricity when solid but does conduct when molten or when dissolved in water. Oxide **P** reacts with water forming a solution with a high pH.

Oxide **Q** is a colourless gas at room temperature. It dissolves in water to give a solution with a low pH.

(i)      Identify **P**. State the type of bonding present in **P** and explain its electrical conductivity. Write an equation for the reaction of **P** with water.

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(ii)     Identify **Q**. State the type of bonding present in **Q** and explain why it is a gas at room temperature. Write an equation for the reaction of **Q** with water.

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**(4)**

(b)     **R** is a hydroxide of a Period 3 element. It is insoluble in water but dissolves in both aqueous sodium hydroxide and aqueous sulphuric acid.

(i)      Give the name used to describe this behaviour of the hydroxide.

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(ii)     Write equations for the reactions occurring.

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(iii)     Suggest why **R** is insoluble in water.

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**(1)**

**(Total 15 marks)**

**Q2.**

Iron(II) sulfate is used to kill weeds in garden lawns. It is a by-product of the manufacture of steel.   
When a lawn is treated with iron(II) sulfate, the iron(II) ions are oxidised to form iron(III) ions.

Iron(III) ions are acidic in aqueous solution as shown by the following equation.

[Fe(H2O)6]3+(aq)   [Fe(H2O)5(OH)]2+(aq) + H+(aq)

In an experiment, a calibrated pH meter was used to measure the pH of an iron(III) salt in solution. At 20 °C the pH of a 0.100 mol dm–3 solution of iron(III) sulfate was found to be 1.62.

(a)     Explain briefly why a pH meter should be calibrated before use.

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(b)     Write an expression for the equilibrium constant, *Ka*, for the dissociation of iron(III) ions in aqueous solution.

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(c)     Use your answer from part (b) to calculate the value of *Ka* for this reaction at 20 °C.   
Give your answer to the appropriate precision. Show your working.

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(d)     Name the substance that is most likely to oxidise the iron(II) ions when iron(II) sulfate is used as a weed killer.

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(e)     Suggest a value for the pH of a 0.100 mol dm–3 solution of iron(II) sulfate.

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**(Total 8 marks)**

**Q3.**

Aqueous metal ions can be identified by test-tube reactions.

          For each of the following, describe what you would observe.

          Write an equation or equations for any reactions that occur.

(a)     The addition of aqueous sodium carbonate to a solution containing  
[Fe(H2O)6]3+(aq) ions.

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(b)     The addition of aqueous sodium hydroxide, dropwise until in excess, to a solution containing [Al(H2O)6]3+(aq) ions.

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**(4)**

(c)     The addition of dilute aqueous ammonia, dropwise until in excess, to a solution containing [Cu(H2O)6]2+(aq) ions.

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(d)     The addition of concentrated hydrochloric acid, dropwise until in excess, to a solution containing [Cu(H2O)6]2+(aq) ions.

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**(2)**

**(Total 14 marks)**

**Q4.**

Which one of the following lists the first ionisation energies (in kJ mol−1) of the elements Mg, Al, Si, P and S in this order?

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **A** | 577 | 786 | 1060 | 1000 | 1260 |
| **B** | 736 | 577 | 786 | 1060 | 1000 |
| **C** | 786 | 1060 | 1000 | 1260 | 1520 |
| **D** | 1060 | 1000 | 1260 | 1520 | 418 |

**(Total 1 mark)**

**Q5.**

In which one of the following reactions does the metal species undergo reduction?

**A**       MnO2 + 4HCl → MnCl2 + 2H2O + Cl2

**B**       [Cu(H2O)6]2++ 4Cl− → [CuCl4]2− + 6H2O

**C**       CrO + 2OH−  → 2CrO + H2O

**D**       TiO2 + 2C + 2Cl2 → TiCl4 + 2CO

**(Total 1 mark)**